

Chemistry – Unit 5 Worksheet 3

Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

1. Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

$$32.0 \text{ g} \times \frac{1 \text{ mole}}{79.9 \text{ g}} = 0.401 \text{ mole Br}$$

1:2 ratio

$$4.9 \text{ g} \times \frac{1 \text{ mole}}{24.31 \text{ g}} = 0.20 \text{ mole Mg}$$

 MgBr_2

2. What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?

$$95.2 \text{ g} - 40.8 \text{ g} = 54.4 \text{ g O}$$

$$40.8 \text{ g} \times \frac{1 \text{ mole}}{12.0 \text{ g}} = 3.40 \text{ moles C}$$

1:1 ratio

$$54.4 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 3.40 \text{ moles O}$$

 CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

$$9.8 \text{ g} \times \frac{1 \text{ mole}}{14.0 \text{ g}} = 0.70 \text{ moles N}$$

1:1:3 ratio

$$0.70 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.69 \text{ moles H}$$

 HNO_3

$$33.6 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 2.10 \text{ moles O}$$

4. A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

$$0.59 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.58 \text{ mole H}$$

1:1 ratio; empirical formula: HO
 empirical formula mass: 17.0 g/mol
 molecular formula mass: 34.0 g/mol
 mfm is 2x efm so the molecular formula is: H_2O_2

$$9.40 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 0.588 \text{ mole O}$$

Empirical Formula Answer Sheet

George Worthington

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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Chemistry Richard Post, Chad Snyder, Clifford C. Houk, 2020-08-20 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of *Chemistry Concepts and Problems A Self Teaching Guide* to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the material Prepare for a career in the sciences medicine or engineering with the core content in this user friendly guide Authored by expert postsecondary educators this unique book gently leads students to deeper levels and concepts with practice critical thinking problem solving and self assessment at every stage

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