

# CH 15 Equilibrium Study Guide

## Dynamic Equilibrium

Dynamic Equilibrium – forward & reverse reaction rates are equal

- **ASSUME** all rxns are **ELEMENTAL** (1 step!)
- Reversible reaction - something constantly produced & consumed
  - Indicated by  $\rightleftharpoons$

No net change in reactant or product concentrations

## Equilibrium Constant, $K_c$

For a general equation:



$$K_c = \frac{k_f}{k_r} = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

- A. The equilibrium-constant depends only on the stoichiometry of the reaction, not the mechanism.  
 B. **DOES NOT** depend on initial concentrations because *only* the **RATIO** of concentrations at equilibrium is the same  
 C. **K** has **NO** units!  
 D. **ONLY** consider gas & aqueous phases!!  
 -OMIT solids & liquids

## Significance of K

Value of K indicates if more products or reactants are present *at equilibrium*

- If K is large (>1): Products favored - Equilibrium lies to the right
- If K is small (<1): Reactants favored - Equilibrium lies to the left

## Equilibrium Constant, $K_p$

For a general equation:

Only if ALL phases are gases!



## Relating $K_p$ and $K_c$

$$K_p = K_c (RT)^{\Delta n} \quad ; \Delta n = [ \sum (\text{moles GAS in product}) - \sum (\text{moles GAS in reactant}) ]$$

# Dynamic Equilibrium Study Guide

**James E. Brady, John R. Holum**



## Dynamic Equilibrium Study Guide:

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