



3.2.6 Study: Dynamic Equilibrium

Physical Science Sem 2

Study Guide

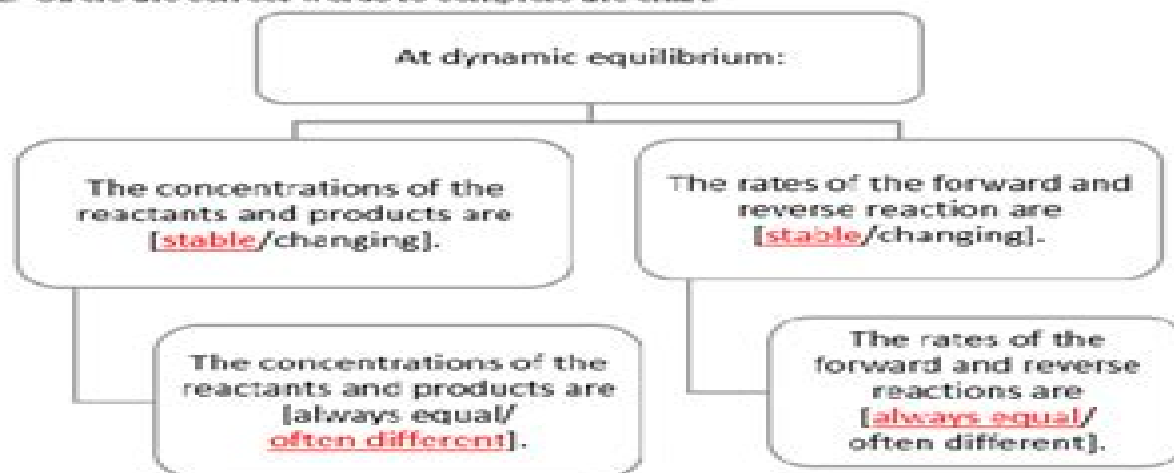
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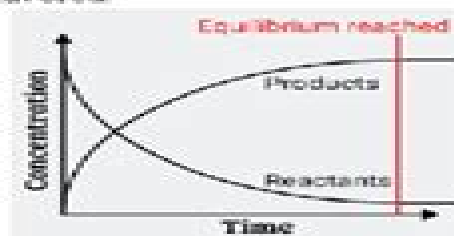
Use this guide to take notes as you work through the activity. Taking good notes can help you remember important ideas. Your notes on this guide will help you prepare for quizzes and tests.

Dynamic Equilibrium

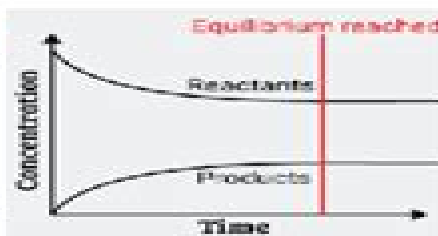
- Fill in the blanks.
 - A reversible reaction can reach dynamic equilibrium because it proceeds in both directions at once.
 - During the forward reaction, the reactants change into products.
 - In the reverse reaction, the products change back into the reactants.
- Circle the correct words to complete the chart.



- Draw and label a line on each graph to show when the reaction reaches equilibrium. Then circle the correct words to indicate whether each reaction is product-favored or reactant-favored.



[product-favored]/reactant-favored] reaction



[Product-favored, reactant-favored] reaction

Le Châtelier's Principle

Dynamic Equilibrium Study Guide

James E. Brady, John R. Holum



Dynamic Equilibrium Study Guide:

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Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **Study Guide for**

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Stovall, Steven S. Zumdahl, 1996 *Chemistry, Study Guide* Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 **Study Guide General Chemistry, Second Edition, Becker/Wentworth** Becker, Benjamin L. Carroll, 1980 [Study Guide and Workbook, an Interactive Approach for Starr and McMillan's Human Biology, Third Edition](#) Jane B. Taylor, John D. Jackson, 1999 Asks the student to write all answers in this study guide workbook This workbook is interactive because it requires students to do things instead of just read more material All questions are arranged by chapter modules so students may skip unassigned material Each module in the study guide refers to the page

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Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

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