

# Chemistry 12

## Solutions Manual Part A

### Unit 2 Structures and Properties of Matter

#### Answers to Unit 2 Preparation Questions (Student textbook pages 158-61)

- Poisonous and infectious material causing immediate and serious toxic effects
  - Use care when handling hot objects
- Tie back long hair. Do not have any flammable material near the burner. Secure the burner. For example, attach it to a retort stand. Handle hot glassware with thermal gloves or beaker tongs.
- The sources must be reliable. Web sites such as university and government sites are most reliable. If the source is posted by an individual, the credentials of the author must be shown to be reliable. Corporate sites are often biased.
- a
- Note to the teacher: The numbers in the first printing of the text were accidentally combined. The values in the shaded squares are the data that should be provided and the open squares are the answers. BLM 3-1 has been provided with the correct working chart.

Name of Isotope	Notation for Isotope	Atomic Number	Mass number	Number of Protons	Number of Electrons
bromine-81	$^{81}_{35}\text{Br}$	35	81	35	35
neon-22	$^{22}_{10}\text{Ne}$	10	22	10	10
calcium-44	$^{44}_{20}\text{Ca}$	20	44	20	20
silver-107	$^{107}_{47}\text{Ag}$	47	107	47	47

- Valence electrons are those electrons that exist in the outer energy level. They are the only electrons that are involved in chemical bonds.
- b
- c
- The first four electrons, in a Lewis diagram, are drawn as far apart as possible because their negative charges repel each other.

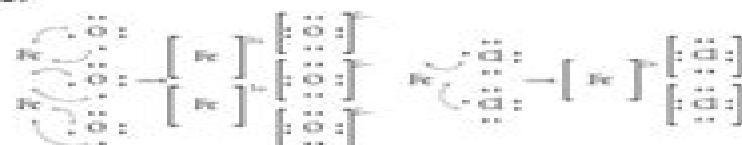
- The distance from the centre of the atom to the boundary within which the electrons spend 90 percent of their time.
  - The energy absorbed or released when an electron is added to a neutral atom.
  - An indicator of the relative ability of an atom to attract shared electrons.
  - The amount of energy required to remove the outermost electron from an atom or ion in the gaseous state.

11.

a. down a group:	b. across a period:
a. increases	a. decreases
b. decreases	b. increases
c. decreases	c. increases
d. decreases	d. increases

- yellow; halogens; very reactive non-metals
  - purple; transition metals; very hard metals with high boiling points
  - brown; other non-metals of the main group elements; most of these atoms are essential building blocks of molecules found in living organisms. Their chemistry typifies how non-metals combine with other atoms.
- In both cases, the metal is donating, in one case  $1e^-$  and in the other  $2e^-$ , to a non-metal. This leaves a positive metal ion and a negative non-metal ion that are attracted to each other by electrostatic forces and they form an ionic compound.
- The octet rule states that when bonds form between atoms, the atoms gain, lose, or share electrons in a way that will create an outer shell containing eight (an octet) of electrons. However, the electron configurations of the transition elements are such that they can have more than eight electrons in their outer shell.

15.



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**Scott C. Dulebohn**



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