

multiply reduction rx. by 3:



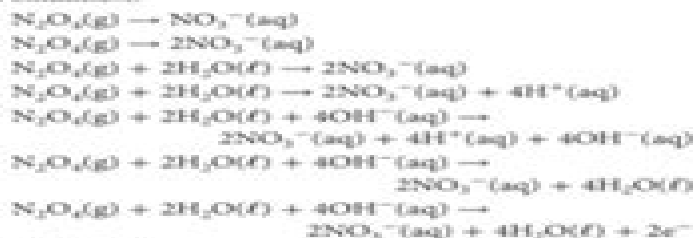
combine:



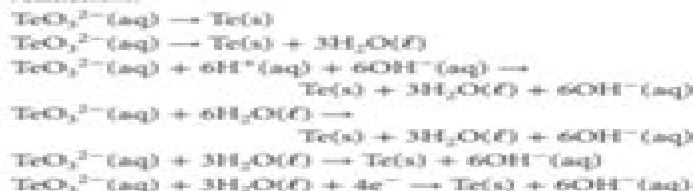
balance:



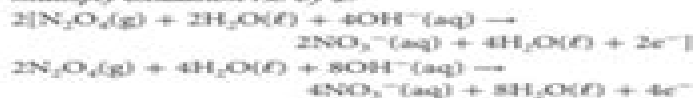
b. oxidation:



reduction:



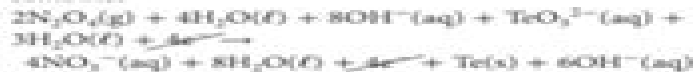
multiply oxidation rx. by 2:



multiply reduction rx. by 1:



combine:



balance:



c. oxidation: $\text{H}_2\text{O}_2(\text{aq}) \rightarrow \text{O}_2(\text{g})$



multiply oxidation rx. by 3:



multiply reduction rx. by 1:



combine:



14. It is a redox reaction because sulfur is both being oxidized from an oxidation number of -2 in $\text{H}_2\text{S}(\text{g})$ to an oxidation number of 0 in $\text{S}(\text{s})$ and reduced from an oxidation number of $+4$ in $\text{SO}_2(\text{g})$ to an oxidation number of 0 in $\text{S}(\text{s})$.

Answers to Section 9.3 Review Questions (Student textbook page 617)

- Fluorine is the most electronegative element in the periodic table. Therefore, it does not "lose" electrons in chemical bonds. It has seven valence electrons, and thus forms bonds in which it gains one valence electron, satisfying the octet rule. Because it is in the second row of the periodic table, its outer occupied energy level cannot accommodate more than eight electrons (as bromine, iodine, or chlorine can). Therefore, it does not form bonds in which it gains more than one electron, and its oxidation number is always -1 .
- MnO_2 $+4$
 - $+7$
 - $+6$
 - KMnO_4 $+7$
- Zero in the diatomic molecule, $\text{O}_2(\text{g})$; -1 in peroxides, $\text{H}_2\text{O}_2(\text{aq})$; -2 in most anions, $\text{CO}_3^{2-}(\text{aq})$; $+2$ in an oxygen difluoride, $\text{OF}_2(\text{g})$
- yes
 - no
 - no
 - yes

Glenco Redox Reaction Solutions Manual

Charles E. Mortimer



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Table of Contents Glenco Redox Reaction Solutions Manual

1. Understanding the eBook Glenco Redox Reaction Solutions Manual
 - The Rise of Digital Reading Glenco Redox Reaction Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Glenco Redox Reaction Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glenco Redox Reaction Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glenco Redox Reaction Solutions Manual
 - Personalized Recommendations
 - Glenco Redox Reaction Solutions Manual User Reviews and Ratings
 - Glenco Redox Reaction Solutions Manual and Bestseller Lists

5. Accessing Glenco Redox Reaction Solutions Manual Free and Paid eBooks
 - Glenco Redox Reaction Solutions Manual Public Domain eBooks
 - Glenco Redox Reaction Solutions Manual eBook Subscription Services
 - Glenco Redox Reaction Solutions Manual Budget-Friendly Options
6. Navigating Glenco Redox Reaction Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Glenco Redox Reaction Solutions Manual Compatibility with Devices
 - Glenco Redox Reaction Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glenco Redox Reaction Solutions Manual
 - Highlighting and Note-Taking Glenco Redox Reaction Solutions Manual
 - Interactive Elements Glenco Redox Reaction Solutions Manual
8. Staying Engaged with Glenco Redox Reaction Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glenco Redox Reaction Solutions Manual
9. Balancing eBooks and Physical Books Glenco Redox Reaction Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glenco Redox Reaction Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Glenco Redox Reaction Solutions Manual
 - Setting Reading Goals Glenco Redox Reaction Solutions Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Glenco Redox Reaction Solutions Manual
 - Fact-Checking eBook Content of Glenco Redox Reaction Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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