

EXPERIMENT 19

ACIDIC AND BASIC SALTS

INTRODUCTION

Water makes up about 60% of the weight of a human body; other animals contain a similar proportion of water, and some plants are almost 95% water. Water covers about 70% of the earth's surface; also, it is in the soil and is present as vapor in the atmosphere. Consequently, reactions that occur in aqueous solutions and reactions in which water participates are very important to us; many are essential to the life and growth of organisms. These reactions might affect the pH of water by using or producing hydronium or hydroxide ions.

It is apparent from the formulas of some compounds that they will produce H_3O^+ or OH^- in aqueous solution. An Arrhenius acid contains hydrogen, which is usually written first in the formula to emphasize its ability to react with water producing H_3O^+ . An ionic Arrhenius base contains hydroxide ions, which separate from the positive ions when the base dissolves in water.

Salts are ionic compounds that contain a cation from a base and an anion from an acid. Even though most salts do not contain H or OH, many of them do react with water changing its pH. Such a reaction is called the *hydrolysis of a salt*, the reaction of an ion of a dissolved salt with water producing H_3O^+ or OH^- .

The hydrolysis of a negative ion produces the negative ion of water. This reaction occurs if the anion is the conjugate base of a weak acid. For example, HCN is a weak acid; $K_a = 4.9 \times 10^{-10}$:



The acid molecule HCN has little tendency to donate protons; its conjugate base, CN^- , is a good proton acceptor. Consequently, when a salt containing CN^- , such as NaCN, dissolves in water, a Brønsted-Lowry acid-base reaction occurs between H_2O (acid) and CN^- (base); that is, a proton is transferred from H_2O to CN^- :



The resulting solution is basic, because hydroxide ions in excess of those produced by the autoionization of water are present.

The hydrolysis of a positive ion produces the positive ion of water. This reaction occurs if the cation is the conjugate acid of a weak base. Molecular bases such as NH_3 are weak bases:



For this reaction, $K_b = 1.8 \times 10^{-5}$. The ammonium ion can donate H^+ to a chemical species that is capable of acting as a base. Consequently, when a salt containing NH_4^+ dissolves in water, NH_4^+ and H_2O react in a Brønsted-Lowry fashion:



Some hydrated metal ions also hydrolyze producing acidic solutions. The strong attraction between a small, highly charged, metal ion and the oxygen atom of a bound water molecule weakens an oxygen-hydrogen bond in the bound water molecule. As a result, the hydrate acts as a proton donor:

Experiment 19 Acids And Basic Salts

Dr. M. Satish Kumar



Experiment 19 Acids And Basic Salts:

An Elementary Experimental Chemistry John Bernard Ekeley, 1900 *Laboratory Experiments in Chemistry to Accompany Black and Conant's "Practical Chemistry"* Newton Henry Black, 1920 *Laboratory Experiments in Chemistry* Newton Henry Black, 1921 **Experiments and Exercises in Basic Chemistry** Steven Murov, Brian Stedjee, 2000-02-03

Practical Physical Chemistry Dr. M. Satish Kumar, Physical Chemistry deals with the relations between the physical properties of substances and their composition The present book is intended to serve as a practical manual for undergraduate and post graduate students I have attempted to assemble the list of experiments from my experience and also have drawn upon the experience of the students who have undergone these laboratory courses and felt the inadequacy of the existing syllabus I am aware that I have not yet exhausted all the experiments that they wanted to place in this book but I had to make a selection keeping the size in consideration This manual is largely structured around the standard experiments of physical chemistry Detailed information on instrumentation kinetics experimental methods and data analysis has been covered I will be happier to take all comments and incorporate them in the further editions **A Laboratory Study of**

Household Chemistry Mary Ethel Jones, 1921 *An Introduction to the study of the compounds of carbon, or, Organic chemistry* Ira Remsen, 1894 *An Introduction to the Study of the Compounds of Carbon* Ira Remsen, 1899 *An*

introduction to the study of the compounds of carbon; or, Organic chemistry. 4th revision Ira Remsen, 1903 **Laboratory**

Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

Chemistry Made Simple John T. Moore, Ed.D., 2010-04-21 See the world one molecule at a time Chemistry helps us understand not only the world around us but also our own bodies CHEMISTRY MADE SIMPLE makes it fun Each chapter has practice problems with complete solutions that reinforce learning A glossary of chemical terms the modern periodic table and detailed illustrations throughout make this the best introduction to one of the most studied of all sciences Topics covered include the Scientific Method the structure and properties of matter compounds laws of chemistry gases liquids and solids solutions electrochemistry the atmosphere biochemistry organic chemistry nuclear chemistry energy the environment Look for these Made Simple titles Accounting Made Simple Arithmetic Made Simple Astronomy Made Simple Biology Made Simple Bookkeeping Made Simple Business Letters Made Simple Earth Science Made Simple English Made Simple French Made Simple German Made Simple Ingles Hecho Facil Investing Made Simple Italian Made Simple Latin Made Simple Learning English Made Simple Mathematics Made Simple The Perfect Business Plan Made Simple Philosophy Made Simple Physics Made Simple Psychology Made Simple Sign Language Made Simple Spelling Made Simple Statistics Made Simple Your Small

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Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section.

An Introduction to the Study of Qualitative Chemical Analysis Launcelot Winchester Andrews, 1891
Syllabus Series University of California (System), 1915 Soil Reaction in Relation to Calcium Adsorption Charles Oscar Swanson, 1922 **Applied Chemistry for Nurses** Joseph Leon Rosenholtz, 1928 **Study Questions and Problems in Inorganic Chemistry** Alexander Silverman, 1924

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