

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A great deal of energy is <u>produced</u> when an ionic compound is formed. |

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Table of Contents Glencoe Chemistry Answer Key Chapter 1

1. Understanding the eBook Glencoe Chemistry Answer Key Chapter 1
 - The Rise of Digital Reading Glencoe Chemistry Answer Key Chapter 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Answer Key Chapter 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glencoe Chemistry Answer Key Chapter 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Answer Key Chapter 1
 - Personalized Recommendations
 - Glencoe Chemistry Answer Key Chapter 1 User Reviews and Ratings
 - Glencoe Chemistry Answer Key Chapter 1 and Bestseller Lists

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 - Glencoe Chemistry Answer Key Chapter 1 Public Domain eBooks
 - Glencoe Chemistry Answer Key Chapter 1 eBook Subscription Services
 - Glencoe Chemistry Answer Key Chapter 1 Budget-Friendly Options
6. Navigating Glencoe Chemistry Answer Key Chapter 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Glencoe Chemistry Answer Key Chapter 1 Compatibility with Devices
 - Glencoe Chemistry Answer Key Chapter 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glencoe Chemistry Answer Key Chapter 1
 - Highlighting and Note-Taking Glencoe Chemistry Answer Key Chapter 1
 - Interactive Elements Glencoe Chemistry Answer Key Chapter 1
8. Staying Engaged with Glencoe Chemistry Answer Key Chapter 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glencoe Chemistry Answer Key Chapter 1
9. Balancing eBooks and Physical Books Glencoe Chemistry Answer Key Chapter 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glencoe Chemistry Answer Key Chapter 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Glencoe Chemistry Answer Key Chapter 1
 - Setting Reading Goals Glencoe Chemistry Answer Key Chapter 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Glencoe Chemistry Answer Key Chapter 1
 - Fact-Checking eBook Content of Glencoe Chemistry Answer Key Chapter 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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