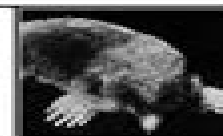


**Activity A:****Moles**Get the Gizmo ready:

- Check that the equation is still
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
- If not, click **New equation** until it reappears.



Introduction: A **mole** is:

- A) A mammal known for digging up gardens.
- B) A small, dark spot on the skin.
- C) A spy embedded within an enemy government.
- D)  $6.02 \times 10^{23}$  particles of a substance.
- E) All of the above.

The correct answer, of course, is E. In chemistry, the mole (mol) is defined as an amount of a substance that contains  $6.022 \times 10^{23}$  particles of that substance. This number, called **Avogadro's number**, is special because this number of particles has a mass in grams that is equal to the mass (in unified mass units) of a single particle of the substance.

Question: How do scientists find the formula mass and molar mass of a substance?

1. 1. **Calculate:** The **formula mass** of a compound is the sum of the masses of each atom in the chemical formula. The unit of formula mass is the unified mass unit (u). Formula mass is also called **molecular mass** if the compound is composed of molecules.

Iron's atomic mass is 55.85 u, carbon's mass is 12.01 u, and oxygen's mass is 16.00 u.

- A. Calculate the formula mass of carbon monoxide (CO) by adding the atomic mass of carbon and the atomic mass of oxygen:

- B. Calculate the formula mass of carbon dioxide (CO<sub>2</sub>):

- C. Calculate the formula mass of iron (III) oxide (Fe<sub>2</sub>O<sub>3</sub>):

2. **Inter:** A mole of a substance has a mass in grams that is equal to the formula mass. For example, a carbon atom has an average mass of 12.01 u. A mole of carbon has a mass of 12.01 g. Based on their formula masses, list the **molar mass** of each substance. The unit for molar mass is g/mol, or grams per mole.

Fe<sub>2</sub>O<sub>3</sub>  CO  Fe  CO<sub>2</sub>

Check your answers on the Gizmo by inspecting the middle row of tiles on the right side of the Gizmo. These tiles show the units "1 mol" on top and "g" below.

3. **Practice:** Hydrogen has an atomic mass of 1.01 u. What is the molar mass of these substances? (Remember to use the units g/mol.)

H<sub>2</sub>O  CH<sub>4</sub>  H<sub>2</sub>CO<sub>3</sub>  C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

# Gizmo Stoichiometry Answers Activity A

**Thomas Griffiths**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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