

# EXPERIMENT: INVESTIGATING PRECIPITATION REACTIONS

PDF

## AIM

To compare the solubility of a range of ionic substances through precipitation.

## EQUIPMENT

☐ test tubes (6)

**Cations** (all 0.1 mol L<sup>-1</sup>) in dropper bottles:

- ☐ sodium nitrate, NaNO<sub>3</sub>
- ☐ silver nitrate, AgNO<sub>3</sub>
- ☐ lead(II) nitrate, Pb(NO<sub>3</sub>)<sub>2</sub>
- ☐ ammonium nitrate, NH<sub>4</sub>NO<sub>3</sub>
- ☐ magnesium sulfate, MgSO<sub>4</sub>
- ☐ iron(II) sulfate, FeSO<sub>4</sub>
- ☐ copper(II) sulfate, CuSO<sub>4</sub>
- ☐ zinc(II) sulfate, ZnSO<sub>4</sub>
- ☐ aluminium sulfate, Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
- ☐ barium chloride, BaCl<sub>2</sub>
- ☐ calcium chloride, CaCl<sub>2</sub>
- ☐ iron(III) chloride, FeCl<sub>3</sub>

☐ test tube rack

**Anions** (all 0.1 mol L<sup>-1</sup>) in dropper bottles:

- ☐ potassium nitrate, KNO<sub>3</sub>
- ☐ potassium chloride, KCl
- ☐ sodium sulfate, Na<sub>2</sub>SO<sub>4</sub>
- ☐ sodium carbonate, Na<sub>2</sub>CO<sub>3</sub>
- ☐ sodium acetate, NaCH<sub>3</sub>COO
- ☐ sodium hydroxide, NaOH



**SAFETY** Wear safety glasses. Lead and barium salts are poisons so avoid contact with skin. If contact occurs wash thoroughly with soap and water. Dispose of as directed by your teacher. Consult "Chemical Safety in Schools" booklet (CDSB).

## PROCEDURE

- 1 Put a 5 mL of a chemical containing each anion into each of six test tubes. Take care not to mix droppers.
- 2 Add 5 mL of a chemical containing the first cation to each. Record your observations.
- 3 Clean all test tubes thoroughly as incorrect results can arise from contamination of solutions.

# Experiment 17 Precipitation Reactions Answers

**Michael Hoelscher**



## Experiment 17 Precipitation Reactions Answers:

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