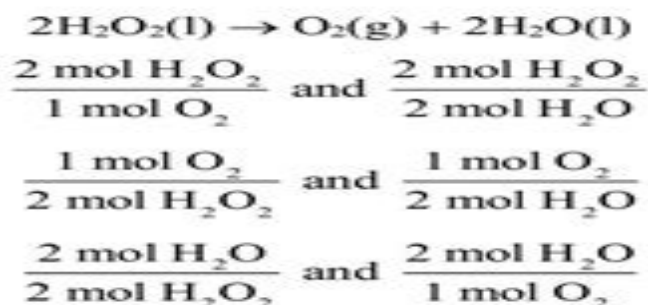


## Chapter 11 (continued)



To determine the number of mole ratios that defines a given chemical reaction, multiply the number of species in the equation by the next lower number. Thus, a chemical reaction with three participating species can be defined by six mole ratios ( $3 \times 2 = 6$ ); a chemical reaction with four species can be defined by 12 mole ratios ( $4 \times 3 = 12$ ).

Why learn to write mole ratios? They are the key to calculations that are based on chemical equations. Using a balanced chemical equation, mole ratios derived from the equation, and a given amount of one of the reactants or products, you can calculate the amount of any other participant in the reaction.

### Practice Problems

2. Determine all the mole ratios for the following balanced chemical equations.
  - a.  $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g})$
  - b.  $4\text{NH}_3(\text{aq}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{l})$
  - c.  $4\text{HCl}(\text{aq}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{Cl}_2(\text{g})$

## 11.2 Stoichiometric Calculations

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There are three basic stoichiometric calculations: mole-to-mole conversions, mole-to-mass conversions, and mass-to-mass conversions. All stoichiometric calculations begin with a balanced equation and mole ratios.

► **Stoichiometric mole-to-mole conversion** How can you determine the number of moles of table salt (NaCl) produced from 0.02 moles of chlorine ( $\text{Cl}_2$ )?

# Exercise 11 Stoichiometry 1

**Raina Robeva**



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