

Isothermal and Adiabatic Processes

Introduction

This experiment measures the pressure, temperature and volume changes in air as it is slowly compressed or expanded under isothermal conditions, or rapidly under near adiabatic conditions.

Theory

The Ideal Gas Law can be written as

$$pV = nRT \quad (1)$$

so that the pressure is inversely proportional to volume when held at constant temperature (isothermal).

An *adiabatic process* takes place when no thermal energy enters or leaves the system. This occurs if the system is perfectly insulated or if the process occurs so rapidly that there is no heat transfer. The first law of thermodynamics for an adiabatic process can be stated as:

$$dQ = nC_v dT + pdV = 0 \quad (2)$$

where C_v is the molar specific heat at constant volume. If we write the ideal gas law as

$$PdV + Vdp = nRdT$$

and solve for dT :

$$dT = \frac{PdV}{nR} + \frac{Vdp}{nR} \quad (3)$$

Famu Physics Lab Manual 2053

David H. Loyd



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