- Caustic soda is obtained on industrial scale by Hg cell which involves the electrolysis of concentrated aqueous solution of sodium chloride using \_\_\_\_\_ anode and \_\_\_\_\_ cathode.
  - a. Ti and Hg
  - b. Hq and Ti
  - c. Hg
  - d. Cu, Ti
- 2. A standard hydrogen electrode has zero electrode potential because:
  - a. It is easy to oxidized hydrogen
  - b. The electrode potential is arbitrarily taken as zero which serves as standard.
  - c. Hydrogen atom has only one electron
  - d. Hydrogen is the lightest element
- The cell in which copper electrode which is dipped in 1M CuSO<sub>4</sub> is connected with SHE through external circuit has the cell potential:
  - a. OV
  - b. +1.10 V
  - c. +0.34V
  - d. -0.34V
- 4. Which of the following element has highest positive standard reduction potential (E°) at 298K?
  - a. Li
  - b. Cu
  - c. Au
  - d. Al
- 5. The element which can displace other halogens from their solutions:
  - a. CI
  - b. F
  - c. Br
  - d. I
- 6. The strongest reducing agent of the alkali metal in solution is:
  - a. Li
  - b. Na
  - c. K
  - d. Cs

# **Electrochemistry Exercise Answers**

**Hussin A.Rothana** 

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