

**AP Chem**  
**Electrochemistry Unit Practice**

1. Tell the oxidation state of each atom in the substance:

- A. Fe
- B.  $O_2$
- C. NaI
- D. BaO
- E.  $MgSO_4$
- F.  $CaH_2$
- G.  $NH_3$
- I. HCN

2. Tell whether each reaction is a redox reaction:

- A.  $NaCl + AgNO_3 = AgCl + NaNO_3$
- B.  $HCl + NaOH = NaCl + H_2O$
- C.  $K + H_2O = KOH + H_2$
- D.  $Cu + 2AgCl = CuCl_2 + 2Ag$

3. Balance each of these half reactions:

- A.  $Cu = Cu^{2+}$
- B.  $Al^{3+} = Al$

4. Balance each of these half reactions in acid:

- A.  $NO_3^- = NO_2^-$
- B.  $ClO^- = ClO_3^-$

5. Balance each of these half reactions in base:

- A.  $Cl_2 = ClO_3^-$
- B.  $MnO_4^- = MnO_2$

6. Determine the number of electrons transferred in these balanced redox equations:

- A.  $MnO_4^- + Cl_2 = Cl^- + Mn^{2+}$
- B.  $HNO_3 + Cu = Cu^{2+} + H_2O + NO_2$

7. Given the two balanced half reactions, write the balanced redox reaction:

- A.  $Li = Li^+ + e^-$ ;  $Zn^{2+} + 2e^- = Zn$
- B.  $Al = Al^{3+} + 3e^-$ ;  $Cu^{2+} + 2e^- = Cu$

8. Balance these equations in acid using the half reaction method:

- A.  $Cr_2O_7^{2-} + NO_2^- = Cr^{3+} + NO_3^-$
- B.  $ClO_3^- + SO_2 = SO_4^{2-} + Cl^-$

9. Balance these equations in base using the half reaction method:

- A.  $Br^- + MnO_4^- = BrO_3^- + MnO_2$
- B.  $H_2O_2 + Cl_2O_7 = ClO_2^- + O_2$

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**Ted Lister, Janet Renshaw**



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reviewing the active and developing aspects of the science of electrochemistry From this beginning has sprung the Electrochemistry Symposium The Past Thirty and the Next Thirty Years which took place at Imperial College London from April 3 6 1975 The plan for this symposium is unusual since it features pairs of invited addresses one to summarize the state of the art and the other to suggest directions for future research in particular aspects of electrochemistry This volume of proceedings gives these papers in their final considered and fully referenced form arranged in the sequence of their delivery at the symposium Also included are introductory addresses given by Professor Ubbelohde Professor Frumkin Dr Egan and Dr Inman Both aspects of nearly every topic plus the discussions are integrated in a Report or Summary A synopsis of the matters raised at the symposium and prepared by Professor John O M Bockris closes this volume The cooperation of Plenum Press New York is gratefully acknowledged

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