

| Substance | Lewis Structure | Molar Mass | Intermolecular Attractive Force |
|---|---|------------|---------------------------------|
| Methanol CH_3OH $12 + 3(1) + 16 + 1 = 32$ | $ \begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \ddot{\text{O}} - \text{H} \\ \\ \text{H} \end{array} $ | 32.042 | Hydrogen Bonds |
| Ethanol $\text{CH}_3\text{CH}_2\text{OH}$ $12 + 5(1) + 16 + 1 = 46$ | $ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array} $ | 46.068 | Hydrogen Bonds |
| n-Butanol $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$ | $ \begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array} $ | 74.122 | Hydrogen Bonds |
| Glycerin $\text{C}_3\text{H}_8(\text{OH})_3$ $12 + 8(1) + 18(16) = 92$ | $ \begin{array}{c} \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \\ \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \end{array} $ | 92.093 | Hydrogen Bonds |
| Water H_2O | $ \begin{array}{c} \text{H} \\ \\ \text{H} - \ddot{\text{O}} - \text{H} \end{array} $ | 18.015 | Hydrogen Bonds |

Exploring Intermolecular Forces Lab Answers

Lei Huang



Exploring Intermolecular Forces Lab Answers:

Exploring General, Organic, & Biochemistry in the Laboratory William G. O'Neal, 2017-02-01 This full color comprehensive affordable manual is appropriate for two semester introductory chemistry courses It is loaded with clearly written exercises critical thinking questions and full color illustrations and photographs providing ample visual support for experiment set up technique and results

Exploring Chemistry (Loose-Leaf) Matthew Johll, 2012-05-04 Matthew Johll's *Exploring Chemistry* covers the standard topics for the nonmajors course in the typical order but each chapter unfolds in the context of a single case study that helps students connect what they are learning to real life situations For example students work through the often difficult topics of molecular structure gas laws and organic chemistry by learning about the development of powerful new chemotherapy drugs new technologies for screening airline passengers and the creation of biodegradable biopolymers It's the same case driven approach that Johll uses in his acclaimed *Investigating Chemistry* now in its Third Edition but *Exploring Chemistry* goes beyond the other book's specific focus on examples from forensic science to use real life stories from cooking athletics genetics green chemistry and more

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Exploring Intermolecular Forces Lab Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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