

Experiment 24

Advance Study Assignment: Molar Mass of an Acid

1. 7.0 mL of 6.0 M NaOH are diluted with water to a volume of 400 mL. You are asked to find the molarity of the resulting solution.

a. First find out how many moles of NaOH there are in 7.0 mL of 6.0 M NaOH. Use Equation 1. Note that the volume must be in liters.

_____ moles

b. Since the total number of moles of NaOH is not changed on dilution, the molarity after dilution can also be found by Equation 1, using the final volume of the solution. Calculate that molarity.

_____ M

2. In an acid-base titration, 23.91 mL of an NaOH solution are needed to neutralize 24.58 mL of a 0.1002 M HCl solution. To find the molarity of the NaOH solution, we can use the following procedure:

a. First note the value of M_{H^+} in the HCl solution.

_____ M

b. Find M_{OH^-} in the NaOH solution. (Use Eq. 3.)

_____ M

c. Obtain M_{NaOH} from M_{OH^-} .

_____ M

3. A 0.3174 g sample of an unknown acid requires 26.23 mL of 0.1056 M NaOH for neutralization to a phenolphthalein end point. There are 0.27 mL of 0.1096 M HCl used for back-titration.

a. How many moles of OH^- are used? How many moles of H^+ from HCl?

_____ moles OH^- _____ moles H^+

b. How many moles of H^+ are there in the solid acid? (Use Eq. 5.)

_____ moles H^+ in solid

c. What is the molar mass of the unknown acid? (Use Eq. 4.)

_____ g/mol

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Yijin Wang



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